

Periodic Trends Lab – Post Lab Questions

1) Which metal reacted faster with water?	2) Which metal reacted faster with acid?
3) Make a statement about the trends in reactivity as you move down the column of alkaline earth metals.	4) Predict the reactivity of strontium and barium, based on your activity in this lab.
5) If sufficient radium could be gathered for a test, predict its reactivity with water and hydrochloric acid. Explain.	6) Why would it be dangerous to handle even a small amount of radium? Your answer should be related to this lab and the concept of reactivity NOT radioactivity.
7) Group VIIA (The halogens – nonmetallic elements) reactivity <i>decreases</i> as the atomic number <i>increases</i> . Why do you think this group of elements is opposite Group IA? Explain. (HINT – think about atomic structure, valence electrons, electronegativity)	
8) Was this a <i>quantitative</i> or <i>qualitative</i> lab? Why?	9) Brainstorm a way that you could add a quantitative aspect to this lab.